


Schedule of Accreditation

issued by

United Kingdom Accreditation Service

2 Pine Trees, Chertsey Lane, Staines-upon-Thames, TW18 3HR, UK

 <p>0658</p> <p>Accredited to ISO/IEC 17025:2017</p>	<p>ROMIL Ltd</p> <p>Issue No: 028 Issue date: 02 October 2025</p>	
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<p>Calibration performed at the above address only</p>		

Calibration and Measurement Capability (CMC)

Measured Quantity Instrument or Gauge	Range	Expanded Measurement Uncertainty ($k = 2$)	Remarks
<p>ASSAY OF CHEMICAL SUBSTANCES FOR USE IN SOLID AND SOLUTION CERTIFIED REFERENCE MATERIALS</p> <p>Adipic Acid Aluminium Nitrate 9H₂O Aluminium metal Ammonium Bromide Ammonium Cerium(IV) Nitrate Ammonium Cerium(IV) Sulphate 2H₂O Ammonium Chloride Ammonium Fluoride Ammonium Hexafluorogermanate Ammonium Hexafluorosilicate Ammonium Iodide Ammonium Nitrate Ammonium Sulphate Ammonium di-Hydrogen Phosphate Antimony Trichloride Antimony Trioxide Arsenic Trioxide Barium Chloride Barium Nitrate Benzoic Acid Bismuth metal Butyric Acid iso-Butyric Acid Boric Acid Cadmium metal Caesium Carbonate Calcium Carbonate Calcium Chloride 2H₂O Calcium Nitrate Cerium(III) Nitrate 6H₂O Chromium(III) Nitrate 9H₂O Chromium(VI) Trioxide Cinnamic Acid Citric Acid 1H₂O Cobalt metal Copper metal Dysprosium(III) Oxide EDTA di-Sodium salt 2H₂O Erbium(III) Oxide</p>	98%-101%	0.10%	Measurement for each by a single primary definitive method at ROMIL using stoichiometric chemical methods



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Measured Quantity Instrument or Gauge	Range	Expanded Measurement Uncertainty ($k = 2$)	Remarks
ASSAY OF CHEMICAL SUBSTANCES FOR USE IN SOLID AND SOLUTION CERTIFIED REFERENCE MATERIALS Europium(III) Oxide Gadolinium(III) Oxide Gallium Nitrate xH ₂ O Glutaric Acid Glycolic Acid Glycine Gold metal Hafnium Dichloride Oxide Holmium(III) Oxide Indium Nitrate 5H ₂ O Iodine Iron(III) Nitrate 9H ₂ O Lactic Acid Lanthanum Nitrate 6H ₂ O Lead metal Lithium Carbonate Lithium-6 Carbonate Lutetium(III) Oxide Magnesium Nitrate 6H ₂ O Magnesium metal Maleic Acid Malic Acid Malonic Acid Manganese metal Mercury(II) Chloride Mercury(II) Nitrate 1H ₂ O Methanesulphonic Acid Molybdenum Trioxide Neodymium(III) Oxide Nickel metal Nitric Acid Oxalic Acid 2H ₂ O Palladium metal Pivalic Acid Potassium Bromide Potassium Carbonate Potassium Chloride Potassium Chromate Potassium Cyanide Potassium Dichromate Potassium Hexacyanoferrate Potassium Hydrogen Carbonate Potassium Hydrogen Phthalate Potassium Iodate Potassium Iodide Potassium Nitrate Potassium Nitrite Potassium Perchlorate Potassium Tetraoxalate 2H ₂ O Potassium Thiocyanate Potassium di-Hydrogen Citrate	98%-101%	0.10%	Measurement for each by a single primary definitive method at ROMIL using stoichiometric chemical methods



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Measured Quantity Instrument or Gauge	Range	Expanded Measurement Uncertainty ($k = 2$)	Remarks
ASSAY OF CHEMICAL SUBSTANCES FOR USE IN SOLID AND SOLUTION CERTIFIED REFERENCE MATERIALS Potassium di-Hydrogen Phosphate Praseodymium(III,IV) Oxide Propionic Acid Rubidium Carbonate Samarium(III) Oxide Scandium Oxide Selenium metal Silver Nitrate Sodium Acetate Sodium Bromate Sodium Carbonate Sodium Chlorate Sodium Chloride Sodium Chlorite Sodium Citrate Sodium Dodecyl Sulphate Sodium Fluoride Sodium Formate Sodium Hydrogen Carbonate di-Sodium Hydrogen Phosphate Sodium Metasilicate 9H ₂ O Sodium Nitrate Sodium Oxalate Sodium Sulphate Sodium Sulphide xH ₂ O di-Sodium Tetraborate 10H ₂ O Strontium Carbonate Succinic Acid Sulphamic Acid Tartaric Acid Tellurium Dioxide Terbium(III,IV) Oxide Thallium Nitrate Thorium Nitrate Thullium(III) Oxide Tin metal Tris(hydroxymethyl)methylamine Vanadium Pentoxide Ytterbium(III) Oxide Yttrium Oxide Zinc Acetate 2H ₂ O Zinc metal Zirconium Dichloride Oxide.xH ₂ O	98%-101%	0.10%	Measurement for each by a single primary definitive method at ROMIL using stoichiometric chemical methods
Other Chemical Substances (The organisation holds a flexible scope for the assay of suitable materials)	98%-101%	0.10%	Measurement for each by a single primary definitive method at ROMIL using stoichiometric chemical methods
pH in buffer solutions	pH 1 to pH 13	0.01 to 0.04 pH units	Measurement by pH meter
END			



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Appendix - Calibration and Measurement Capabilities

Introduction

The definitive statement of the accreditation status of a calibration laboratory is the Accreditation Certificate and the associated Schedule of Accreditation. This Schedule of Accreditation is a critical document, as it defines the measurement capabilities, ranges and boundaries of the calibration activities for which the organisation holds accreditation.

Calibration and Measurement Capabilities (CMCs)

The capabilities provided by accredited calibration laboratories are described by the Calibration and Measurement Capability (CMC), which expresses the lowest measurement uncertainty that can be achieved during a calibration. If a particular device under calibration itself contributes significantly to the uncertainty (for example, if it has limited resolution or exhibits significant non-repeatability) then the uncertainty quoted on a calibration certificate will be increased to account for such factors.

The CMC is normally used to describe the uncertainty that appears in an accredited calibration laboratory's schedule of accreditation and is the uncertainty for which the laboratory has been accredited using the procedure that was the subject of assessment. The measurement uncertainty is calculated according to the procedures given in the GUM and is normally stated as an expanded uncertainty at a coverage probability of 95 %, which usually requires the use of a coverage factor of $k = 2$. An accredited laboratory is not permitted to quote an uncertainty that is smaller than the published measurement uncertainty in certificates issued under its accreditation.

Expression of CMCs - symbols and units

It should be noted that the percentage symbol (%) represents the number 0.01. In cases where the measurement uncertainty is stated as a percentage, this is to be interpreted as meaning percentage of the measurand. Thus, for example, a measurement uncertainty of 1.5 % means $1.5 \times 0.01 \times q$, where q is the quantity value.

The notation $Q[a, b]$ stands for the root-sum-square of the terms between brackets: $Q[a, b] = [a^2 + b^2]^{1/2}$